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**Highly sensitive and selective Love mode surface acoustic wave ammonia  
sensor based on graphene oxides operated at room-temperature**

Q. B. Tang<sup>a</sup>, Y. J. Guo<sup>a,\*</sup>, Y. L. Tang<sup>b</sup>, G. D. Long<sup>a</sup>, J. L. Wang<sup>a</sup>, D. J. Li<sup>c</sup>, X. T.  
Zu<sup>c</sup>, J. Y. Ma<sup>d</sup>, L. Wang<sup>d</sup>, H. Torun<sup>e</sup>, Y. Q. Fu<sup>e,\*</sup>

<sup>a</sup> School of Physics, University of Electronic Science and Technology of China,  
Chengdu, 610054, P. R. China

<sup>b</sup> School of Physical Science and Technology, Southwest Jiaotong University,  
Chengdu, 610031, P. R. China

<sup>c</sup> Institute of Fundamental and Frontier Sciences, University of Electronic  
Science and Technology of China, Chengdu, 610054, P. R. China

<sup>d</sup> Sichuan Institute of Piezoelectric and Acousto-Optic Technology, Chongqing,  
400060, P. R. China

<sup>e</sup> Faculty of Engineering and Environment, University of Northumbria,  
Newcastle upon Tyne, NE1 8ST, UK

**Abstract:** It is **crucial** to develop highly sensitive and selective sensors for ammonia, one of the most common toxic gases which have been widely used in **pharmaceutical, chemical and manufacturing** industries. In this study, graphene oxide (GO) film was spin-coated onto surfaces of ST-cut quartz surface acoustic wave (SAW) devices with a resonant frequency of 200 MHz for ammonia sensing. The oxygen-containing functional groups (**such as hydroxyl and epoxy ones**) on the **surface of** GO film **strongly** absorb ammonia molecules **and thus increase the film stiffness**. This is attributed to the main **ammonia sensing** mechanism **of** the Love mode SAW devices, **which show not only** a positive frequency shift of 620 Hz for 500 ppb ammonia gas, **but** also an excellent selectivity (as compared to other gases such as H<sub>2</sub>, H<sub>2</sub>S, CO and NO<sub>2</sub>) **and** a good reproducibility, **operated at room temperature of 22 °C**.

**Keywords:** Surface acoustic wave (SAW); graphene oxide (GO); ammonia sensor.

\*Corresponding authors.

*E-mail address:*

Prof. Yuanjun Guo ([guoyuanjun@uestc.edu.cn](mailto:guoyuanjun@uestc.edu.cn)); Prof. Richard Yongqing Fu ([richard.fu@northumbria.ac.uk](mailto:richard.fu@northumbria.ac.uk))

## 1. Introduction

Ammonia, one of the most common toxic gases, is widely used in pharmaceutical, chemical and manufacturing industries. Ammonia is considered to be a very promising hydrogen carrier due to its high energy density, low cost, easy liquefaction which can be transported easily. In addition, ammonia is essential for carbon-free production of hydrogen [1, 2]. Due to its wide-range of applications, the risk for ammonia leakage or contamination in our daily life becomes quite high, and excessive inhalation of ammonia can lead to respiratory disease, severe headache, sore throat, loss of smell and chest pain. Occupational Safety and Health Administration, United States of Labor set a permissible exposure limit for ammonia at 50 ppm during a work day of eight hours. Therefore, it is critical to develop highly sensitive and highly selective ammonia sensors, and significant attention has recently been paid on detection and monitoring of ammonia along with other toxic gases [3-5].

Surface acoustic wave (SAW) technology has been well developed for gas sensing and particularly for ammonia sensing [6-13]. SAW devices generally have high sensitivity, fast response, good stability, and they offer ease of subsequent signal processing. The operation principle of a SAW gas sensor is based on the interaction of a sensitive film with the targeted gas molecules that results in changes in phase velocity and amplitude of the waves. Therefore, the concentration of the target gas can be measured by monitoring the changes in frequency or phase angle of the resonant frequency peaks [14]. Adding a sensitive layer material which has a strong affinity to

the targeted gas is an effective way to improve the sensitivity [15]. A variety of materials have been employed for the sensitive layers, including polymers [16-18], semiconducting metal oxides [7, 11, 19, 20], carbon nanomaterials [21, 22], and graphene and its derivatives [23, 24]. When these layers are deposited onto a quartz-based shear-horizontal SAW device, a Love mode SAW device can be developed, which has significantly improved the sensitivity due to the existence of the guiding layer for the SAWs [13].

Graphene derivatives (such as graphene oxide: GO, and reduced graphene oxide, rGO) have been reported to have larger surface areas and more chemically active sites than those of nanostructured metal oxides, conductive polymers, and carbon nanotubes [25]. In addition to their intrinsic advantages, their chemical and electrical properties are tunable based on their growth parameters. These characteristics are very attractive for gas sensing applications. GO, a precursor to graphene, contains a variety of oxygen-containing functional groups. It is generally accepted that these functional oxygen groups exist mainly in the form of hydroxyl groups and epoxy groups on the surface of the GO sheet, together with smaller amounts of carboxyl groups, carbonyl groups, quinone, lactones, and phenols at the edges of the GO sheet [26-28]. Due to the presence of these oxygen-containing groups, GO can be stably dispersed in many solvents, especially in water [29-31]. A uniform layer of GO can be obtained on various substrates by using simple methods such as dip coating, spray coating or spin coating [27]. In addition, GO possesses the 2D characteristics of graphene, which could

enhance the effective interactions with many analytes.

Various GO-based sensors have been reported for gas monitoring. GO-based SAW sensors have been demonstrated for the detections of humidity [24], NO<sub>2</sub> [32, 33], H<sub>2</sub> [34] and chemical warfare agents [35]. One of the advantages of using the GO for ammonia detection was demonstrated by Peng et al. [36] based on first-principles calculations, which revealed that the GO can be used to detect ammonia more efficiently than graphene. GO's surface-active defect sites such as epoxy and hydroxyl groups promote the interactions between ammonia molecules and GO. The GO has been widely used in optical [37] and impedance sensors [36] to detect ammonia. Generally, the impedance-based GO sensors are often needed to operate at elevated temperatures (e.g. often above 100°C) in order to increase their responses and recovery rates and sensitivities [38, 39]. However, when the operating temperature is higher than 200 °C, the oxygen-containing functional groups are decomposed, leading to an apparent degradation in sensitivity [40].

A good technique for the realization of precision sensing of ammonia operated at room temperature is SAW technology, especially those Love-mode SAW devices. Moreover, SAW sensors can be operated wirelessly with a low power consumption and a capability of remote operation. This is practically advantageous for toxic gas sensors. However, currently there are only a limited number of studies for the GO-based SAW sensors for the ammonia [41, 42], thus it is extremely valuable to prepare a SAW sensor based on the GO for ammonia sensing and study its sensing mechanisms.

In this work, we realized a GO-based Love mode SAW ammonia sensor on an ST cut quartz substrate, operated at a room temperature of 22 °C. Its key sensing mechanism is mainly based on the increase of the sensing layer's stiffness when the ammonia molecules are adsorbed by the oxygen-containing functional group of GO, which are deposited onto the SAW surface. Our results indicate the presence of oxygen-containing functional groups also enhances sensitivity and selectivity of the sensor operated at room temperature.

## 2. Experimental methods

The two-port resonator SAW device was fabricated on an ST-cut quartz substrate using the conventional photolithography and metallization processes. Aluminum layer with a thickness of 200 nm was sputtered onto the ST-cut (42°75') quartz substrate and then the interdigital transducers (IDT) were prepared using a lift-off process. The propagation direction of the generated SH-SAW is perpendicular to the crystallographic x-axis (90°-rotated). There are two sets of 30 pairs of IDTs on the SAW resonator. The period of IDT is 16 μm and the aperture is 3 mm. The center frequency of SAW devices was designed to be ~200 MHz, which is determined by the wave velocity on the piezoelectric substrate (3158 m/s) and the SAW wavelength (16 μm). The SAW devices had a center frequency of 200.02 MHz and an insertion loss of -8.82 dB, as shown in Figure 1. GO was produced using an improved Hummers' method [43]. The prepared GO was dissolved in deionized water, and then sonicated for one hour to prepare an aqueous solution with 2 mg/ml graphene oxide. The prepared GO solution was spin-

coated onto the SAW substrate at a spin-speed of 6000 rpm for 30 seconds, and then dried at 60 °C for 10 minutes, and then the Love mode SAW device was obtained. The above processing cycle was repeated for five times to obtain a ~50 nm-thick GO sensing film. Our results showed that when the GO film thickness exceeds 50 nm, the current test setup does not work properly due to the limited gain of the amplifier circuit.

Field-emission scanning electron microscope (FE-SEM, FEI Inspect F50) was used to estimate the size of the prepared GO platelet and characterize the surface morphology of the prepared films. In addition, the GO film was characterized using various methods, including Raman spectroscopy (Alpha 300R with a laser beam wavelength of 488 nm), Fourier transform infrared spectroscopy (FTIR, Nicolet IS 10, in the range of 400–4000  $\text{cm}^{-1}$ ) and X-ray photoemission spectroscopy (XPS, Kratos Axis Ultra DLD, Al  $K\alpha$  radiation ).

The resonant frequency of the SAW resonators was measured using a frequency counter (Agilent 53210A) with the help of a peripheral matching circuit. The SAW sensor was placed in a testing chamber with a volume of 20 L, which was pre-flowed with dry nitrogen gas before introducing ammonia. The standard purchased gas of ammonia (2%  $\text{NH}_3$  with the balanced gas of  $\text{N}_2$ ) was purchased from the National Institute of Testing Technology, China. The ammonia concentration in the testing chamber was precisely controlled using a dynamic volume method, e.g., 10 ml of standard ammonia gas was mixed with air (with a fixed ambient relative humidity, RH, of 40%) to obtain 10 ppm of ammonia gas in the gas chamber, and different ammonia



concentrations were obtained by controlling the volume of ammonia injected inside the chamber. During the gas sensing tests, the ambient temperature was controlled at 22°C and the ambient RH value was fixed at 40%. We monitored the resonant frequency and the electrical parameters of the Love-mode SAW device at different ammonia concentrations using a vector network analyzer (Hewlett Packard 8714C). In addition, the sheet resistance of the GO sensitive film and its changes after ammonia injection were measured using a source meter (Keithley 2400).

### **3. Results and discussions**

#### **3.1 Characterization of sensing film**

An aqueous solution of 0.1 mg/ml of GO was coated on a silicon wafer and this GO layer was observed using the FE-SEM. The image is shown in Fig. 2, which shows that the prepared GO platelets have sizes ranging from 0.5 to 4  $\mu\text{m}$ . Fig. 3 shows the SEM images of GO film, exhibiting a continuous membrane with some creases, and the inset image shows the cross-sectional SEM image, revealing that the thickness of the film is about 50 nm.

Fig. 4(a) shows the Raman spectrum of the GO film, with the D peak at 1356  $\text{cm}^{-1}$  and G peak at 1581  $\text{cm}^{-1}$ . The former corresponds to the structural defects caused by the hydroxyl and epoxy groups on the carbon surface [44], and the latter corresponds to the bond stretching of all carbon  $\text{sp}^2$  components on the rings and chains [40, 45]. Other peaks at 120  $\text{cm}^{-1}$ , 204  $\text{cm}^{-1}$ , and 458  $\text{cm}^{-1}$  are from the quartz substrate

[46]. FTIR spectra of the GO film is shown in Fig. 4(b), in which the oxygen-containing groups with different modes of vibrations can be detected, including epoxides (C-O-C, 1415 and  $\sim 918\text{ cm}^{-1}$ ), carboxyl (COOH,  $1599\text{ cm}^{-1}$ ), basal hydroxy and hydroxy (edges Hydroxyl) (C-OH,  $3100\text{-}3600\text{ cm}^{-1}$ ). Both the Raman and FTIR results clearly showed that there are significant amount of epoxy, carboxyl, and hydroxy groups on the surface of the GO.

Fig. 5 shows XPS results of the GO film. Fig. 5(a) shows the XPS survey spectrum, from which the C/O atomic ratios of GO were measured to be 2.25. Figs. 5(b) and 5(c) are the C1s and O1s spectra of the GO film, respectively. The C1s spectrum consists of three well-resolved binding energy configurations at 284.9, 286.8, and 288.2 eV, corresponding to bonds of C-C, C-O and C=O, respectively. The C-C bonds reflect graphitized  $\text{sp}^2$  hybridized carbon. The C-O bonds are attributed to the epoxy group and the hydroxyl group and phenol located on the surface of the GO. The C=O bonds are attributed to the carboxyl group and carbonyl group at the edge of the GO or at the defect. Figure 5(b) shows the O1s spectrum with the binding energy range of 528-536 eV, which can be deconvoluted into C=O (531.6 eV), C-O-C (532.1 eV) and C-OH (533.7 eV). The C=O, C-O-C and C-OH bonds are corresponding to those of carbonyl/carboxy group, epoxy group and hydroxyl group on the GO, respectively. These results are consistent with those in the literature, indicating the presence of carbonyl/carboxy, epoxy and hydroxyl groups on surface of the GO [47].

### 3.2 Gas sensing performance and sensing mechanisms

Fig. 6(a) shows the dynamic responses of the Love mode sensor when it is exposed to ammonia concentrations between 0.5 ppm and 100 ppm at the room temperature of 22 °C and an RH of 40%. Our experimental data suggest that the resonant frequency increases with the concentration of the ammonia and the sensor can detect 500 ppb ammonia with an increase of the resonant frequency up to 620 Hz. It is also observed that the sensor's response increases with increasing the ammonia concentration. Fig. 6(b) shows that the frequency shift increases drastically at low concentrations of ammonia, whereas the reaction sites are gradually saturated at high concentrations due to the existence of limited functional groups that result in a modest frequency shift. If the frequency shift is re-drawn as a function of logarithm of ammonia concentration, a linear relationship can be obtained as shown in Fig. 6(c), which can be expressed using the following equation:

$$\Delta f = 1601 + 3904 \times \lg(C_{\text{ppm}}) \text{ (Hz)} \quad (1)$$

where  $\Delta f$  is the frequency shift in Hz and the  $C_{\text{ppm}}$  is the  $\text{NH}_3$  concentration in ppm.

Fig. 6(d) shows the response and recovery times at different concentrations. Here, the response time is defined as the time required for the device's frequency shift to reach 90% of its final value. Similarly, the recovery time is defined as the time required for the device's frequency shift to return back to 10% of its total frequency shift value after the target gas is released. The recorded response times of the sensor exhibit some variations, but are all less than 250 s, whereas the recovery times are longer, with values

between 400 to 700 s. Table 1 summarizes the detection limits of various ammonia SAW sensors reported in the literature, and it is clear that the Love-mode SAW sensor we prepared shows the lowest detection limit.

Changes in the resonant frequency of the Love mode SAW devices during gas exposure can be attributed to various phenomena such as mass loading, acoustoelectric effect (electrical loading), and elastic loading [15].

The changes in the electrical properties (in particular, conductivity and dielectric constant) of a thin film is due to the adsorption and reaction of target gas molecules. **Because** the **GO** sensing layer has significant amount of epoxy, carboxyl, and hydroxy groups on **its** surface, **and** these will result in changes in the speed of sound, known as electroacoustic effect [15]. When the targeted gas molecules are introduced into the sensor surface, the frequency shift ( $\Delta f$ ) will be positive if the conductivity of the sensing layer is decreased. Otherwise, the shift will be negative. The relationship between  $\Delta f$  and electroacoustic effect can be represented using the following equation [15]:

$$\frac{\Delta f}{f_0} = \frac{\Delta v}{v_0} \approx -\frac{k^2}{2} \left( \frac{\sigma_s^2}{\sigma_s^2 + v_0^2 c_s^2} \right) = -\frac{k^2}{2} \left( \frac{1}{1 + \left( \frac{v_0 c_s}{\sigma_s} \right)^2} \right) \quad (2)$$

where  $f_0$  (~200 MHz) is the intrinsic resonant frequency;  $k^2$  (0.11% for quartz) is the electromechanical coupling factor;  $v_0$  (3158 m/s for quartz) is the unperturbed wave velocity for the SAW device; and  $c_s$  (0.5 pF cm<sup>-1</sup>) is the sum of the dielectric permittivity of the substrate and the environment.

For the Love mode SAW sensor, the acoustoelectric effects depend on the changes of surface conductivity ( $\sigma_s$ ) and resistance. Fig. 7(a) shows that the changes of the electrical conductivity of sensing layer will lead to frequency changes. As shown in Fig. 7(a), when the acoustoelectric parameters ( $\xi = \sigma_s / v_0 C_s$ , in which  $V_0$  is 3158 m/s and  $C_s$  is 0.5 pF/cm) is between 0.01 and 100, an increase in the conductivity of the film causes a negative frequency shift, while a decrease in the conductivity of the film results in a positive frequency shift [15]. The sheet conductivity of GO film  $\sigma_s$  in the air is  $5.4 \times 10^{-4} \text{ Sm}^{-1}$ , and the calculated acoustoelectric parameter is 340. The calculated values of the frequency shift due to the conductivity changes as shown in Fig. 7(a) are very small as compared to the experimental data which we obtained from this study. For example, assuming that the conductivity of the GO film is reduced to one percent of the previous one after injecting 10 ppm of ammonia gas, the calculated frequency shift caused by the acoustoelectric effect is only 9.6 Hz.

In order to further prove this conclusion experimentally, we measured the voltage vs. current results of the SAW electrode when the GO film was exposed to 500 ppm ammonia. We then calculated the resistance changes of the electrode, which are linked with the conductivity changes. The parameter for the resistance response of the SAW device was defined as  $R = R_{\text{air}}/R_{\text{gas}}$ , in which  $R_{\text{air}}$  is the resistance of GO film in ambient air and the  $R_{\text{gas}}$  is the resistance of GO film in ammonia and air mixed atmosphere. The calculated resistance response values of the SAW device are shown in Fig. 7(b) without any significant changes (i.e, varied from 1.0034 to 0.9985). Based on our theoretical

analysis and experimental data, we conclude that the electroacoustic effect is not dominant for the sensor's main frequency shift.

The mass loading on the film changes the resonant frequency of sensors as expressed in equation (3) [51]:

$$\frac{\Delta f}{f_0} = (k_1 + k_2) f_0^2 \rho_s \quad (3)$$

where  $k_1 = -8.7 \times 10^{-8} \text{ m}^2 \text{ s kg}^{-1}$  and  $k_2 = -3.9 \times 10^{-9} \text{ m}^2 \text{ s kg}^{-1}$  are material constants of the ST-cut quartz substrate;  $f_0$  is the center frequency of the SAW that is not disturbed ( $\sim 200 \text{ MHz}$ );  $\rho_s$  is the surface density of the layer. When the mass density of the GO film is increased in ammonia gas, it will result in a negative frequency shift, and if its density is reduced, a positive frequency shift will occur.

When the GO is exposed to ammonia gas, the hydroxyl, carboxyl and epoxy groups on the sensitive film have a complex adsorption process with ammonia molecules with the help of water in the air [52-56]. The GO's adsorption of ammonia can be divided into physical adsorption and reactive adsorption. If reactive adsorption occurs (such as carboxyl and epoxy groups form amines after reacting with ammonia), the sensor will not be fully recovered and the repeatability will be poor [54,55]. Our experimental results show that the GO SAW ammonia sensor can be fully recovered, so we believe that the GO's adsorption of ammonia in this study is only a physical adsorption. When the GO layer physically adsorbs the ammonia gas molecules, the density of the sensitive layer will be increased and the resonance frequency shift will be decreased, thus resulting in a negative frequency shift. However, we have

systematically observed an increase in resonant frequency with the increase of gas concentration. Therefore, we believe the mass loading is not the key sensing mechanism for our sensor.

Another major influence for the frequency shift of the SAW device is the elastic loading [57]. When the sensing membrane adsorbs molecules, its viscoelasticity changes, and the generated stress or strain can partially be transmitted to the substrate. If there is a densification effect, the resonant frequency will increase. The elastic loading could be the dominant mechanism for the increase in the resonant frequency that we observed in this study. The frequency change due to the elastic loading can be expressed using the following equation[58]:

$$\frac{\Delta f}{f_0} = c_e f_0 h \left( \left( \frac{4u}{v_0^2} \right) \times \left( \frac{u + \lambda}{u + 2\lambda} \right) \right) \quad (4)$$

where  $c_e$  is the sensitivity coefficients of elasticity,  $f_0$  is the unperturbed frequency,  $h$  is film thickness,  $v_0$  is the unperturbed SAW velocity,  $\lambda$  is the bulk modulus of elasticity, and  $u$  is the shear modulus of elasticity. When the value of  $u$  is small ( $\approx 10^6 Pa$ ), elastic loading can be negligible. For a large value of shear modulus ( $\approx 10^9 Pa$ ), the elastic loading becomes critical and can induce a positive frequency shift [59].

According to previous reports [60],  $\lambda$  and  $u$  can be expressed by Young's modulus and Poisson's ratio as follows:

$$u = \frac{E}{2(1 + \nu)} \quad (5)$$

and

$$\lambda = \frac{E}{2(1+\nu)} \quad (6)$$

where  $E$  and  $\nu$  are the Young's modulus and Poisson's ratio of the materials, respectively. Therefore, we can rewrite Eq. (3) as follows [57]:

$$\frac{\Delta f}{f_0} = c_e f_0 h \left( \left( \frac{4u}{v_0^2} \right) E \times \left( \frac{4\nu - 5}{5\nu^2 + \nu - 4} \right) \right) \quad (7)$$

GO has a large Young's modulus, tens of GPa [61,62]. FTIR and XPS results in Section 3.1 show that different hydroxyl, carboxyl and epoxy groups exist on the surface of GO film. The hydroxyl groups and epoxy groups on GO film have shown large adsorption energies for ammonia molecules [36], indicating that GO has good adsorption to ammonia. Due to the network structure of GO and the small size of ammonia molecules, the ammonia molecules could easily be filled within the gaps between different GO atomic layers, thus increasing the stiffness of film, and resulting in a positive frequency shift, which is the same with the experimental result.

We further tested the selectivity of our GO SAW sensor against various major gas targets. Fig. 9(a) shows the sensor's dynamic responses to different gases. The concentrations of H<sub>2</sub>, H<sub>2</sub>S, CO and NO<sub>2</sub> are 50 ppm each, whereas that of NH<sub>3</sub> is 20 ppm. The GO SAW sensor is insensitive to H<sub>2</sub> and CO gases. It also shows minor responsivity to NO<sub>2</sub> and H<sub>2</sub>S gases with the measured frequency shift values being less than 600 Hz. Whereas this GO SAW sensor shows a significant response to ammonia gas with measured frequency shift values of as large as 6700 Hz. Results showed that



our GO SAW sensor has an excellent selectivity for  $\text{NH}_3$  gas.

Peng et al. reported that the oxygen-containing functional groups of GO (such as hydroxyl groups and epoxy groups) have good adsorption capability to ammonia [36].

Ammonia is highly soluble in water (volume ratio of ammonia and water is 1:700), so the surface of GO can effectively absorb both moisture and ammonia. In addition, ammonia incorporated by GO forms a new adsorption center to enhance the adsorption of ammonia [53]. Ammonia adsorption by the GO layers increases the stiffness of the sensitive layer, resulting in a large positive shift, thus explaining its excellent selectivity.

Finally, we tested the repeatability of our Love mode SAW sensor to the ammonia gas. Fig. 9(b) shows the measured dynamic responses of our sensor to 20 ppm  $\text{NH}_3$  exposure for four successive cycles. The measured frequency shifts are 6.86 kHz, 7.01 kHz, 7.09 kHz, 7.33 kHz, respectively, after each of the four cycles. The variation in the resonant frequency is less than 5%, indicating the good repeatability of the SAW sensor.

#### 4. Conclusions

We have developed the Love mode SAW resonators coated with GO sensing film on ST-cut quartz substrates for the detection of ammonia at room temperature of 22 °C. The sensor exhibits good responses to  $\text{NH}_3$  gas, and it can detect 500 ppb  $\text{NH}_3$  gas exposure with a positive frequency shift of 620 Hz. In addition, we characterized the relevant sensing parameters of our device. We measured response times less than 250 s and recovery times less than 700 s at room temperature. We also observed excellent

selectivity against major gas targets of H<sub>2</sub>, H<sub>2</sub>S, CO, NO<sub>2</sub> and good reproducibility at room temperature. In addition to demonstration of the sensing performance, we have systematically investigated the dominant sensing mechanism of our sensor by evaluating mass loading effect, acoustoelectric effect, and elastic loading effect. For this purpose, we characterized the structure of the sensor using XPS, Raman, and FTIR analyses. Results confirm the presence of different hydroxyl groups, carboxyl groups, and epoxy on the GO film. When exposed to ammonia, these oxygen-containing functional groups adsorb ammonia molecules, and increase the stiffness of the film, which is identified as the main reason for the positive frequency shift of the sensor.

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## FIGURE CAPTIONS

Figure 1. Transmission signal  $S_{21}$  of the ST-cut quartz SAW resonators.

Figure 2. Optical microscope image showing the feature of GO platelet (marked with a circle)

Figure 3. Top-view and cross-view SEM images of GO film.

Figure 4. (a) Raman spectra of GO film showing the existence of D peak and G peak; (b) FTIR spectra of GO film showing the existence of epoxides, carboxyl, hydroxyl and phenolic groups.

Figure 5. XPS spectra of GO film. (a) high resolution XPS spectrum of C 1s; (b) high resolution XPS spectrum of O 1s.

Figure 6. (a) Dynamic frequency responses of the GO SAW sensor exposed to different concentrations of  $\text{NH}_3$  gas; (b) frequency shift of the sensor exposed to different  $\text{NH}_3$  concentrations; (c) frequency shift with a log function of ammonia concentration; (d) response and recovery times of the sensor at different  $\text{NH}_3$  concentrations.

Figure 7. (a) Relationship between normalized frequency shifts with different acoustoelectric parameters when exposed to the ammonia gases with different concentrations; (b) Electrical resistance responses ( $R=R_{\text{air}}/R_{\text{gas}}$ ) of the GO film when exposed to 500 ppm  $\text{NH}_3$ .

Figure 8. The reaction relationship between GO oxygen-containing functional groups and ammonia after adsorption of  $\text{NH}_3$ .

Figure 9. (a) Dynamic responses of the SAW sensor when exposed to different types of gases; (b) Dynamic responses of the sensor to  $\text{NH}_3$  gas of 20 ppm for successive four cycles.

Fig.1

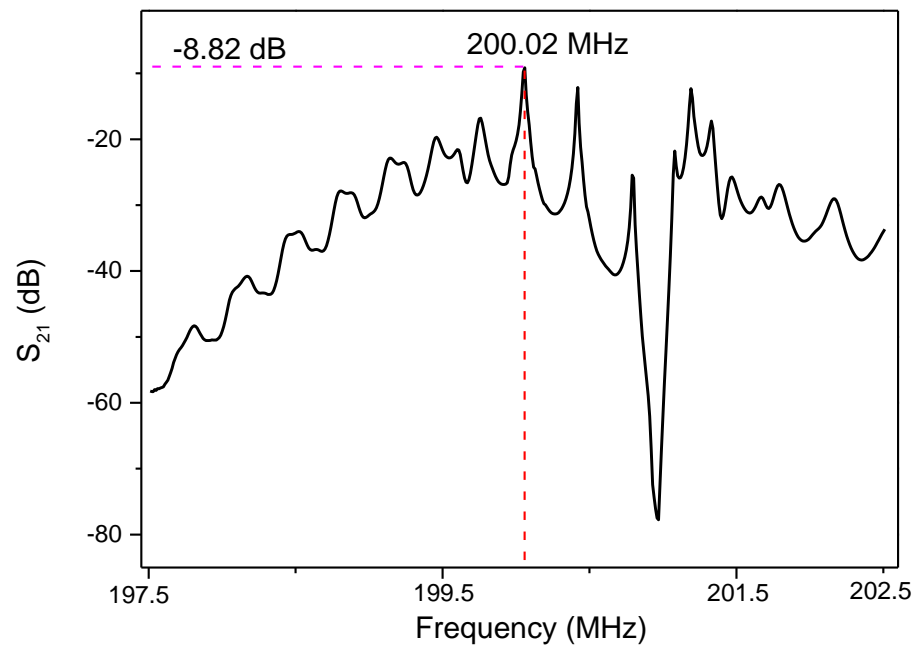




Fig.2

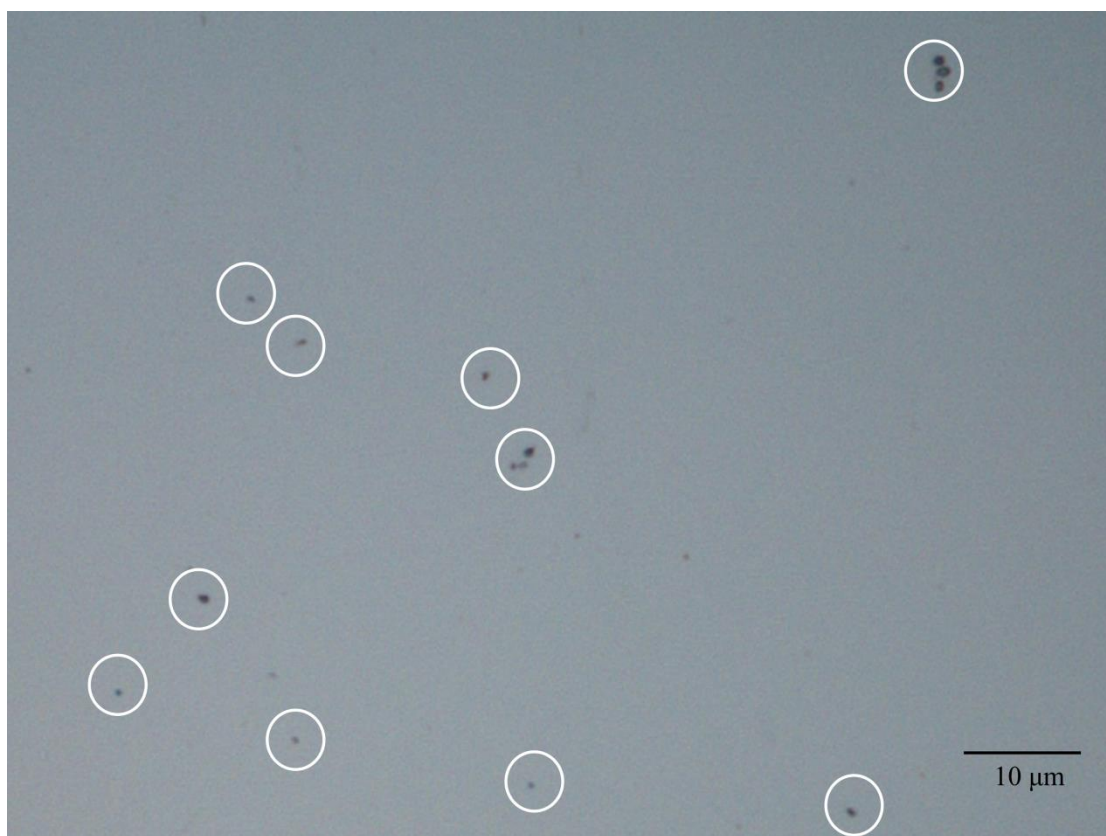


Fig.3

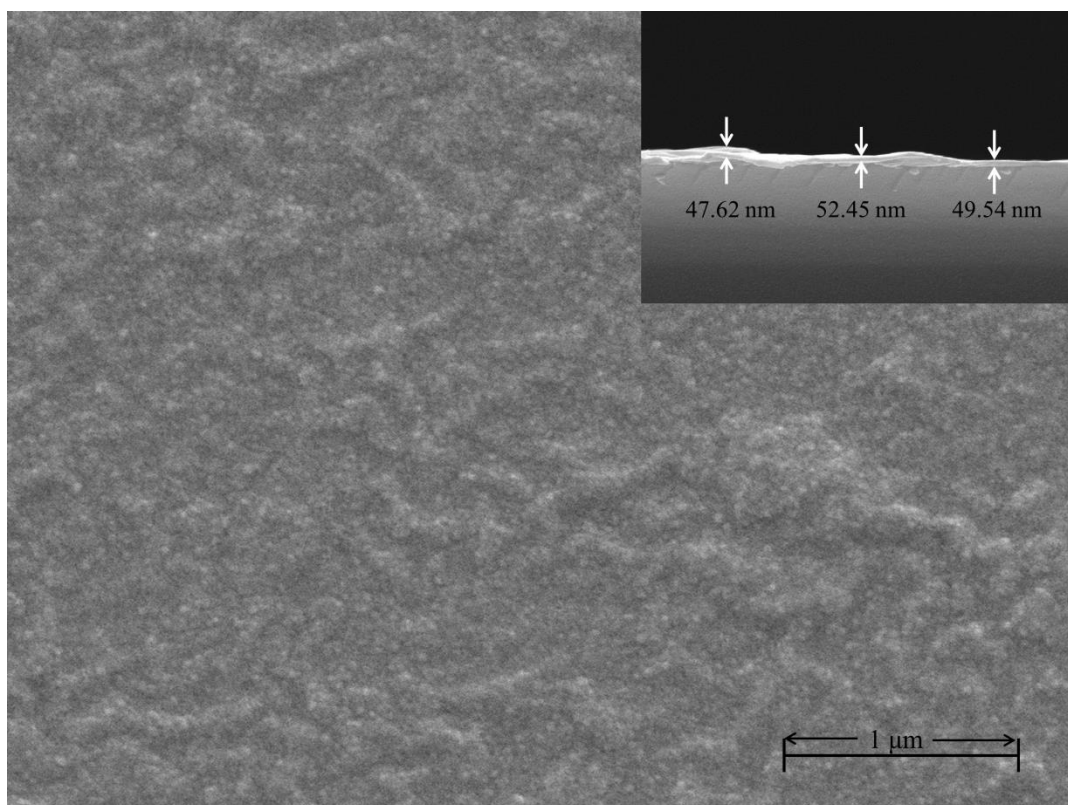


Fig.4(a)

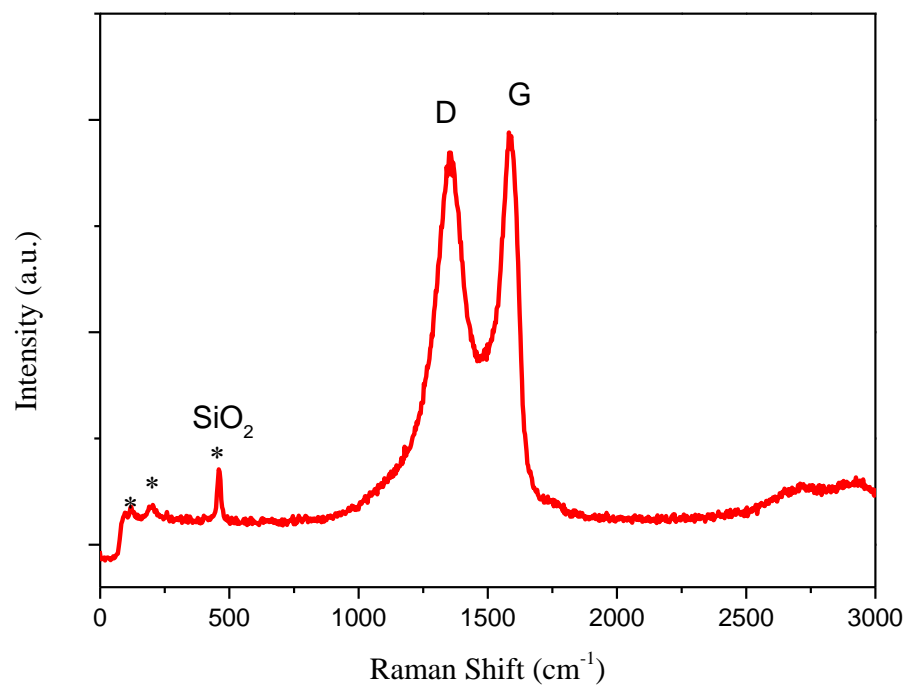


Fig.4(b)

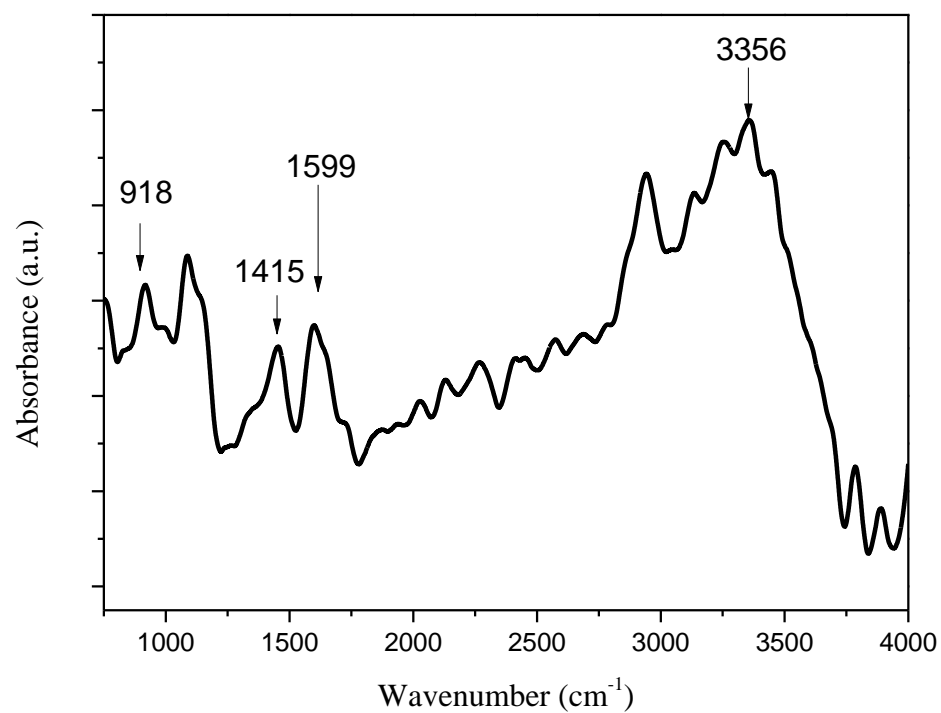


Fig.5(a)

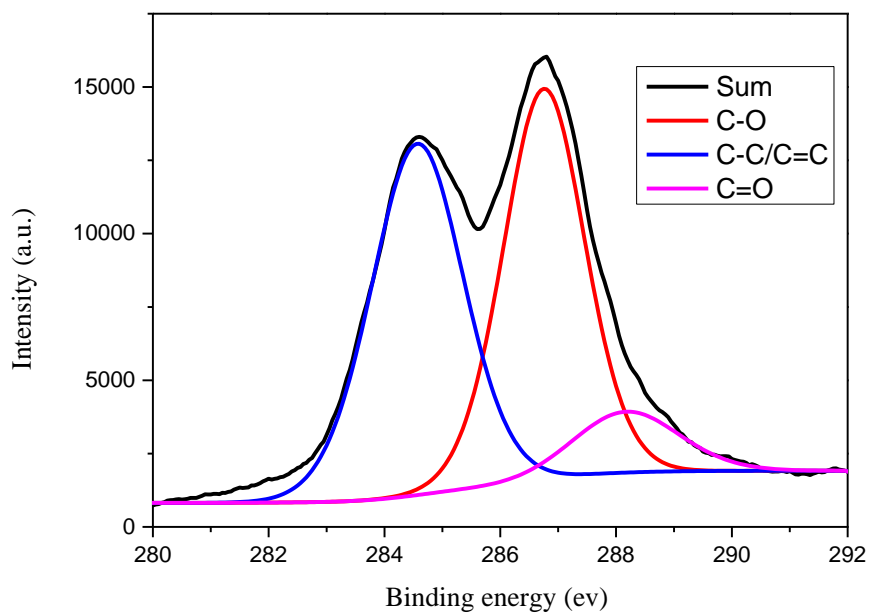


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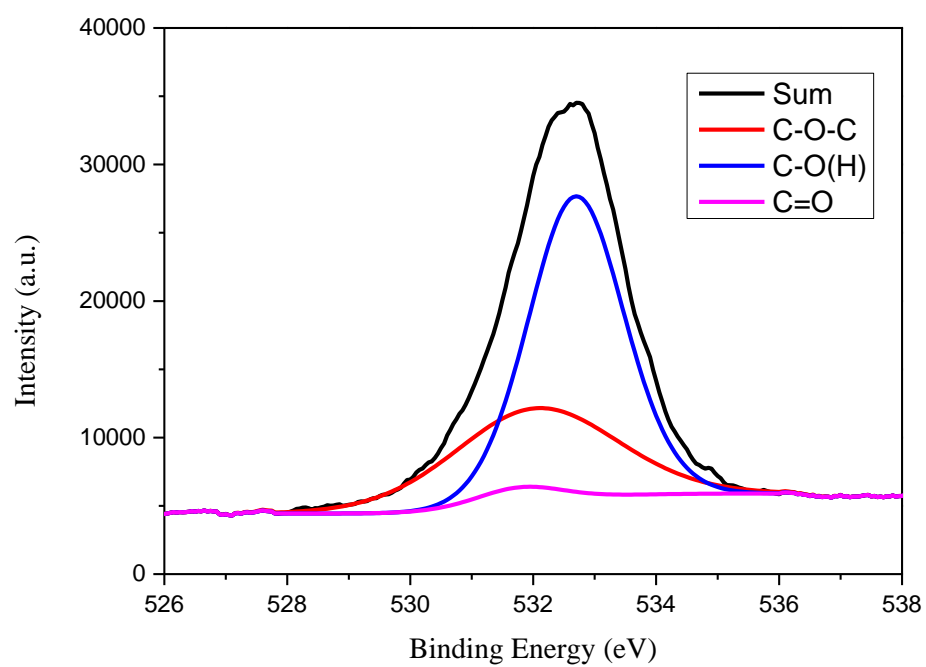


Fig.6(a)

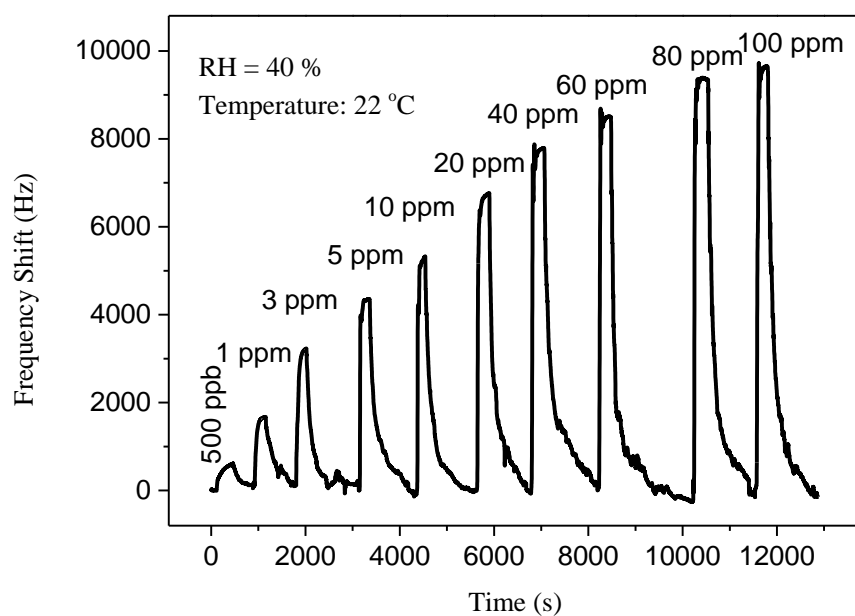


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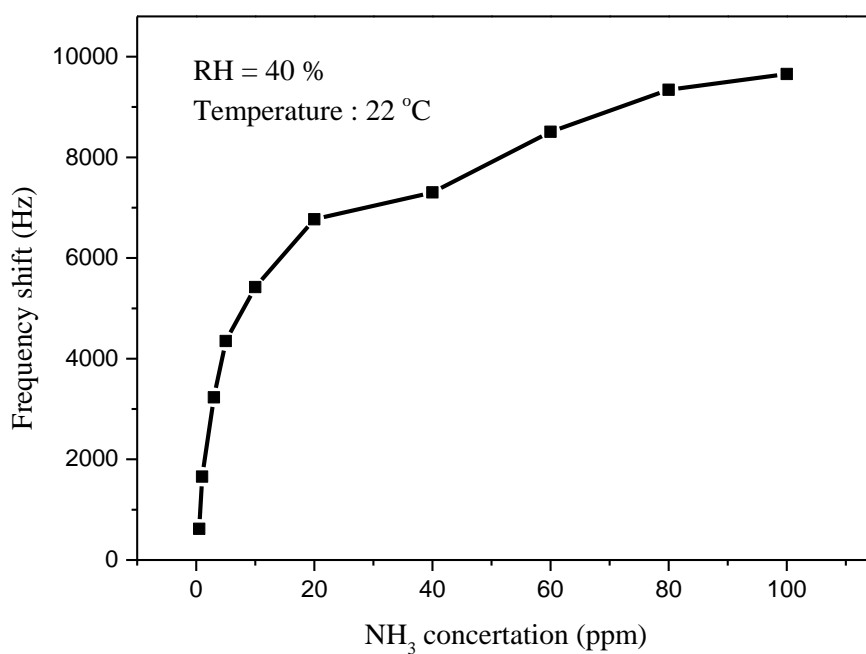




Fig.6(c)

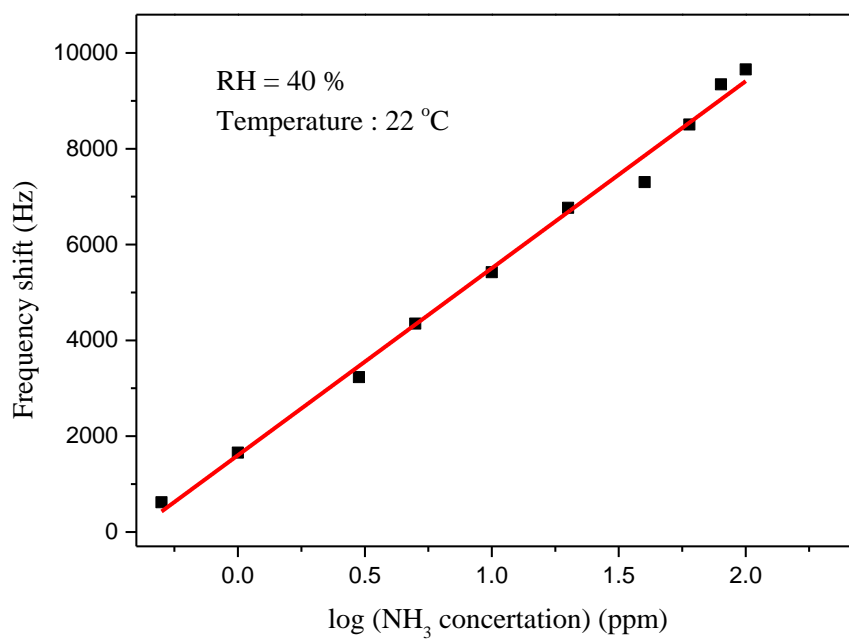


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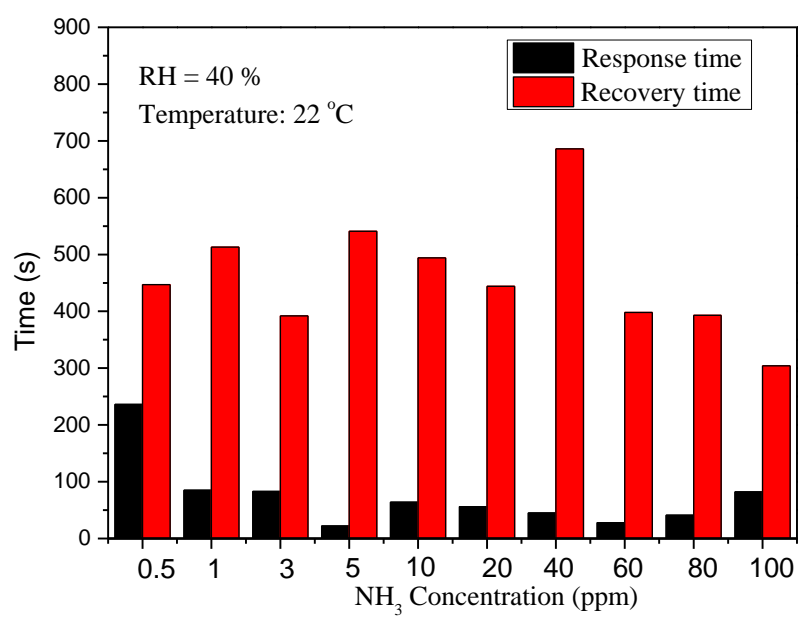


Fig.7(a)

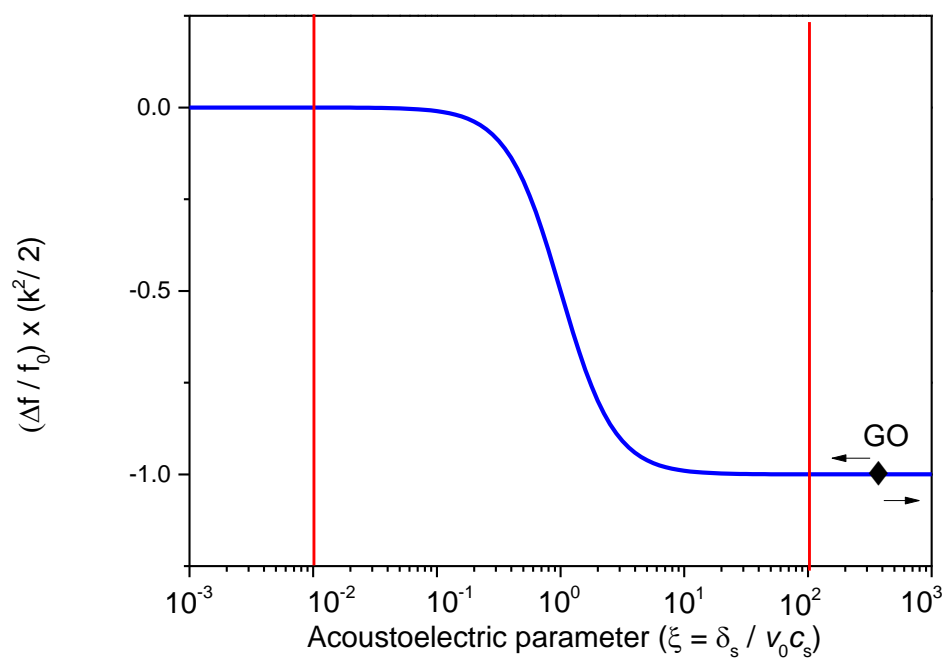


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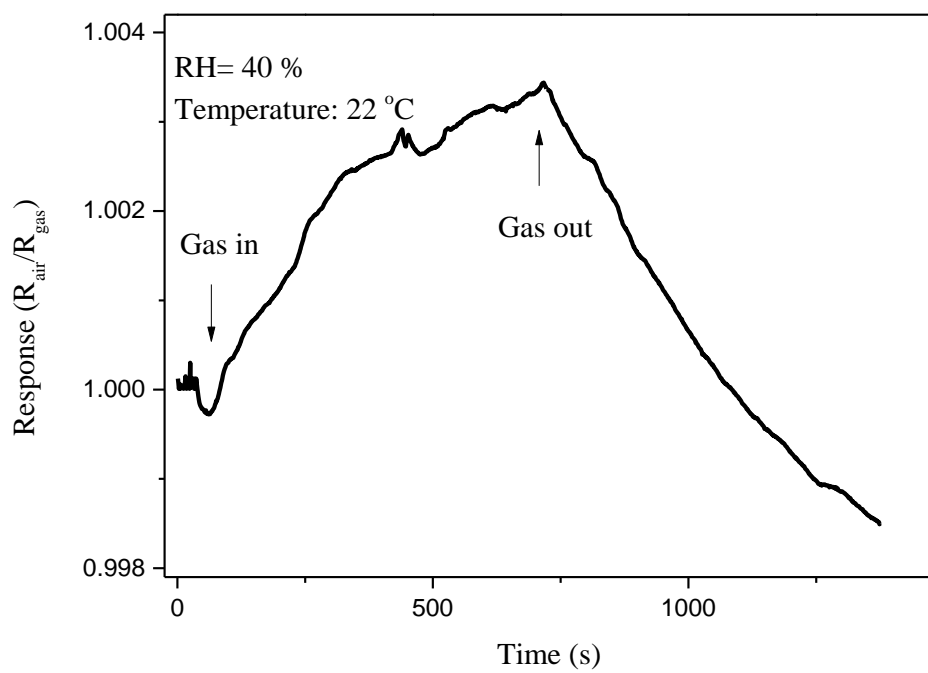


Fig.8

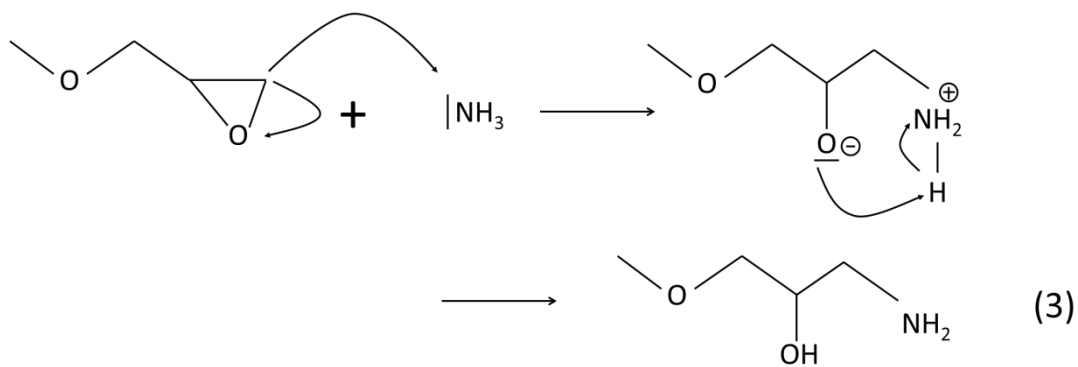
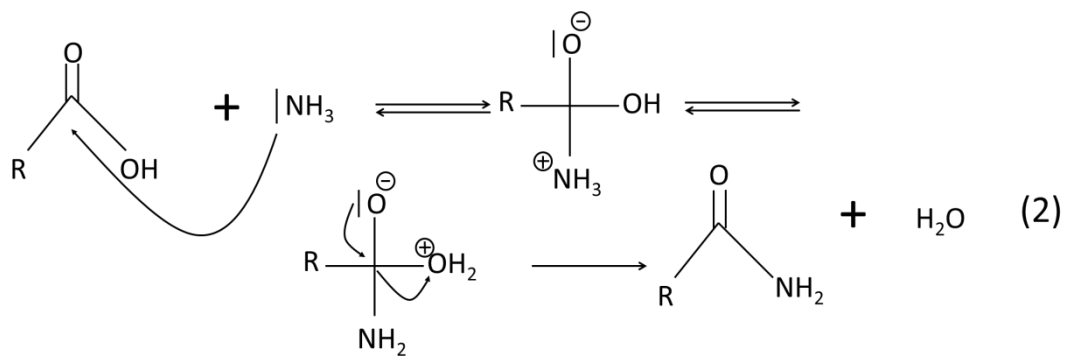
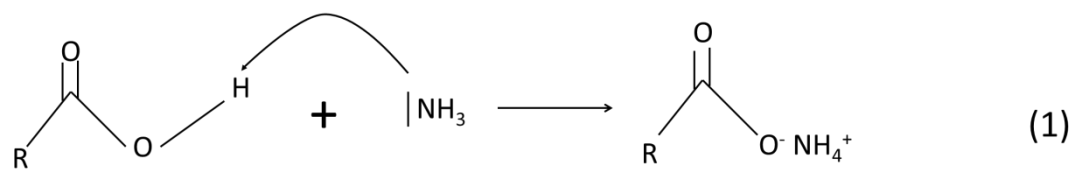


Fig.9(a)

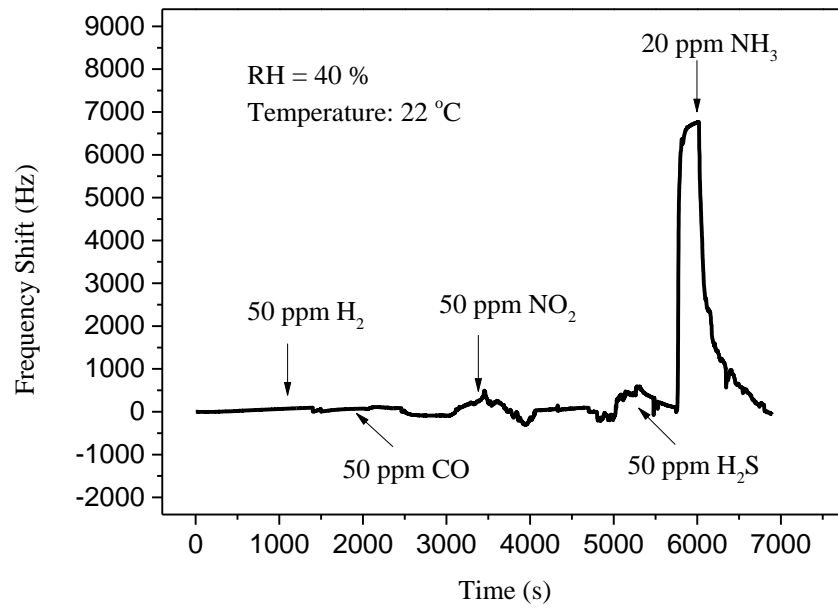


Fig.9(b)

